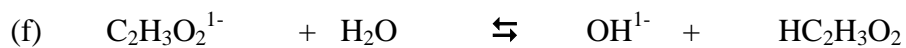
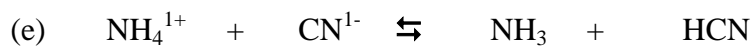
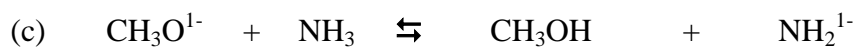
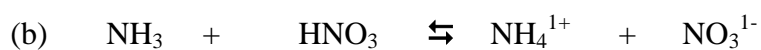


Acids & Bases Worksheet #1
(2009 Version)

In each of the following equations, label the acid, base, conjugate acid, and conjugate base.



The Mathematics of Acids and Bases and the pH Scale

$$\text{pH} = -\log [\text{H}^{1+}]$$

$$\text{pOH} = -\log [\text{OH}^{1-}]$$

$$\text{pH} + \text{pOH} = 14.00$$

$$[\text{H}^{1+}] [\text{OH}^{1-}] = 1.0 \times 10^{-14}$$

pH Scale: 0 – 14

pH < 7.0
ACIDIC

pH = 7.0
NEUTRAL

pH > 7.0
BASIC

Complete the table by calculating the missing entries and indicating if the solutions are acidic or basic.

[H ⁺]	[OH ⁻]	pH	pOH	Acidic or Basic?
0.0032				
	0.000067			
		8.25		
			5.70	
		4.22		
0.55				
			8.91	
	0.100			
		11.90		
	0.00250			
0.00015				
			2.77	